

# PHYSICAL PHARMACY

**PHYSICAL CHEMISTRY** – branch of chemistry concerned with physical properties and structure of matter and with laws and theories of physical and chemical changes

**PHYSICAL PHARMACY** – application of physical chemistry in pharmacy; study of physicochemical properties of substances used in drug formulation

## I. FORCES OF ATTRACTION

### A. INTRAMOLECULAR – intra (within) the molecule

- **Ionic Bond** – transfer of electrons; metal (electron donor) + non-metal (electron acceptor); (ex. Na:  $p^+ = 11$ ;  $e^- = 11 \rightarrow p^+ = 11$ ;  $e^- = 10 \rightarrow$  charge = +1)

#### ○ PERIODIC TRENDS REPRESENTING BONDS

- **Ionization Energy** – energy required remove electron (low M; high NM)
- **Electronegativity** – ability to attract electrons (low M; high NM)

- **Covalent Bonds** – sharing of electrons; non-metal + non-metal; polarity – difference in electronegativity (symmetrical = non-polar)

- **Polar** – unequal sharing of electrons; dipole moment – (ex.  $H^{\delta+}-Br^{\delta-}$ )
- **Non Polar** – equal sharing of electrons (ex.  $Br-Br$ )

### B. INTERMOLECULAR – inter (between) the molecules

- **Van der Waals**

- **Dipole-Dipole** → P + P; Keesom Forces; orientation effect
- **Dipole-Induced Dipole** → P + NP; Debye Forces; inductive effect
- **Induced Dipole-Induced Dipole** → NP + NP; London Forces; dispersive effect

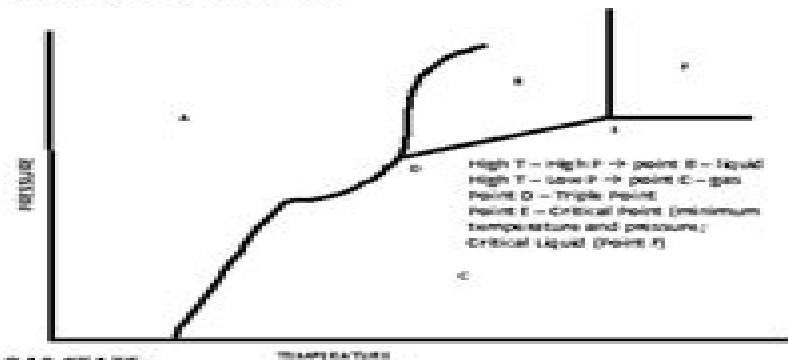
- **Ion-Induced Dipole** → ion + NP

- $I_2 \rightarrow NP$  (solubility = 1:2960)
- $I_2 + NaI/KI$  (solubilizing agent) →  $I_3 + I_2$  = increase solubility

- **Hydrogen Bond** – hydrogen atom + electronegative atom (N, O, F)



## PHASE DIAGRAM



### GAS STATE:

- **Kinetic Molecular Theory**
  - Gases are composed of small particles called atoms; total volume is negligible in relation to space where they are confined
  - Particles do not attract one another but move with complete independence
  - Particles exhibit continuous random motion owing to their kinetic energy
  - Gases exhibit perfect elasticity (no energy exchange)

## II. GAS LAWS

- at standard temperature (Kelvin =  $^{\circ}C + 273$ ) and pressure (1 atm = 760 mmHg or torr; 760 mmHg;  $1.01325 \times 10^5$  dynes/cm<sup>2</sup>;  $1.01325 \times 10^5$  Pa or N/m<sup>2</sup>; 0.1325 bar)

- **Boyle's Law** – constant temperature; increase P = decrease V; inversely related

$$P \propto \frac{1}{V}$$

- \* density = increase altitude; decrease volume;  $P_1V_1 = P_2V_2$

- **Charles' Law** – constant pressure; increase T = increase V

$$T \propto V; \frac{T_1}{T_2} = \frac{V_1}{V_2} \text{ or } \frac{T_1}{V_1} = \frac{T_2}{V_2}$$

- \*  $H_2O$  – 1.0g/ml → 4 $^{\circ}C$

- **Gay Lussac's Law** → constant volume

$$P \propto T; \frac{P_1}{P_2} = \frac{T_1}{T_2} \text{ or } \frac{P_1}{T_1} = \frac{P_2}{T_2}$$

- \* 121 $^{\circ}C$  – 15-20psi

# Physical Pharmacy Lecture Notes

**David Kirk**



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