

**Physical Science – 2nd Semester – Final Exam Study Guide – (2012–2013)****1. Physical Science Basics**

- a. What tool(s) would you want to use to find the:
- Mass of an object? **triple beam balance** Basic SI Unit? **kg**
  - Volume of an object? **ruler or water displacement method** Basic SI Unit? **cm<sup>3</sup> or mL**
  - Length of an object? **Ruler** Basic SI Unit? **meter**
  - Density of an object? **graduated cylinder** Basic SI Unit? **mL**
- b. What is the formula for calculating density? **mass/volume**
- What is the density of a cube that is 33 g and has a volume of 3 cm<sup>3</sup>? **11g/cm<sup>3</sup>**
  - Compare the density of solid H<sub>2</sub>O to liquid H<sub>2</sub>O, and explain what observable effect this has. **solid H<sub>2</sub>O less dense than liquid H<sub>2</sub>O**

**2. Chemistry Basics**

- a. Element (def.) **substance made up of one kind of atom**
- b. Atom (def.) **basic building block of matter**
- c. Compound (def.) **a molecule that contains atoms or molecules of two or more different elements**
- d. Molecule (def.) **a combination of two or more of the same or different atoms that are bonded together**

**3. Model of the Atomic Evolution of the Atomic Theory**

Atoms are so small that many scientists throughout history have created **models** to describe them.

Why have atomic theories changed over time? **better technology came around.**

Scientist	Year	Major Discovery/ Contribution	Name of atomic model
John Dalton	1808	Each element is made up of small, indivisible particles called atoms. Different elements have atoms of different masses.	Billiard Ball Model
J.J. Thomson	1897	Negatively charged electron.	Plum Pudding Model
Niels Bohr	1913	Electrons move <b>fixed orbits</b> around the atom's nucleus in specific layers, or <b>electron shells</b> . Atoms absorb or give off energy when the electrons move from one shell ( <b>energy level</b> ) to another.	Bohr Model
Ernest Rutherford	1909	1909 (Based on idea that positive reflects positive) atom mostly empty space; electrons orbit randomly around positively charged nucleus.	Planetary Model

**4. Periodic Table**

- a. The current periodic table is arranged by increasing **atomic number**.
- b. Reading the Periodic Table
- An element's **properties** can be predicted from its location in the periodic table.
  - The atomic number represents the number of **protons**.
  - The atomic mass represents the number of **protons and neutrons**.

# Physical Science Final Exam 3 Study Guide

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web dec 12 2022 at any time the picu will be home to patients with a variety of diagnoses such as patients with respiratory disorders including acute asthma exacerbations pneumonia burns and other trauma seizure disorders toxic ingestions seasonal respiratory viruses such as rsv and other serious infections

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web nurse practitioners the mott picu nurse practitioner program began in 2006 our very own longstanding picu bedside nurse janella reske was the founding member the program has grown its ranks to eleven recruiting nursing expertise from diverse acute care settings the nurse practitioners are integrated team members practicing

**advanced roles on picu paediatric critical care society** - Apr 11 2023

web advanced nurse practitioner s answers the gap in the career prospects for picu nurses wanting to focus on the clinical domain advanced nurse practitioners within picu undertake a role that is clinically focused with responsibilities in all four pillars of advanced practice clinical practice education leadership and research

**regulation of pediatric intensive care unit nurse practitioner practice** - Jun 13 2023

web purpose to describe the extent to which organizational regulation of pediatric intensive care unit picu nurse practitioner np practice and prescriptive authority aligns with state scope of practice ssop regulations to examine differences between picu medical directors and nps report of regulation and to describe organizational level