

Calculate the final pH of a buffer containing 0.1 M sodium acetate and 0.2 M acetic acid ( $pK_a = 4.76$ ) following the addition of 0.05 M NaOH.

$$pH = pK_a + \log \frac{\text{Base}}{\text{Acid}}$$

$$pH = 4.76 + \log \frac{0.1 + 0.05}{0.2 - 0.05}$$

$$pH = 4.76 + \log \frac{0.15}{0.15}$$

$$pH = 4.76$$

# Ph Properties Of Buffer Solutions Lab Calculations

**Guillaume Favre**



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