

Class 10th Chemistry Practical Copy

Lab No. _____ Page No. _____

A solution can be prepared only by combining the hydrogen ion concentration of an acid with the concentration of metal cations and an anion part of it. Salt has been prepared with metal cation part of it. The pH value was found to be _____

Acid solution: $[H^+] = 10^{-2} M$, pH = 2.00
 Basic solution: $[OH^-] = 10^{-2} M$, pH = 12.00
 Neutral solution: $[H^+] = [OH^-] = 10^{-7} M$, pH = 7.00

1. Solution having a pH of 2 will be a neutral solution with pH value above 7 and below 7 and solution with pH value below 7 are acids.

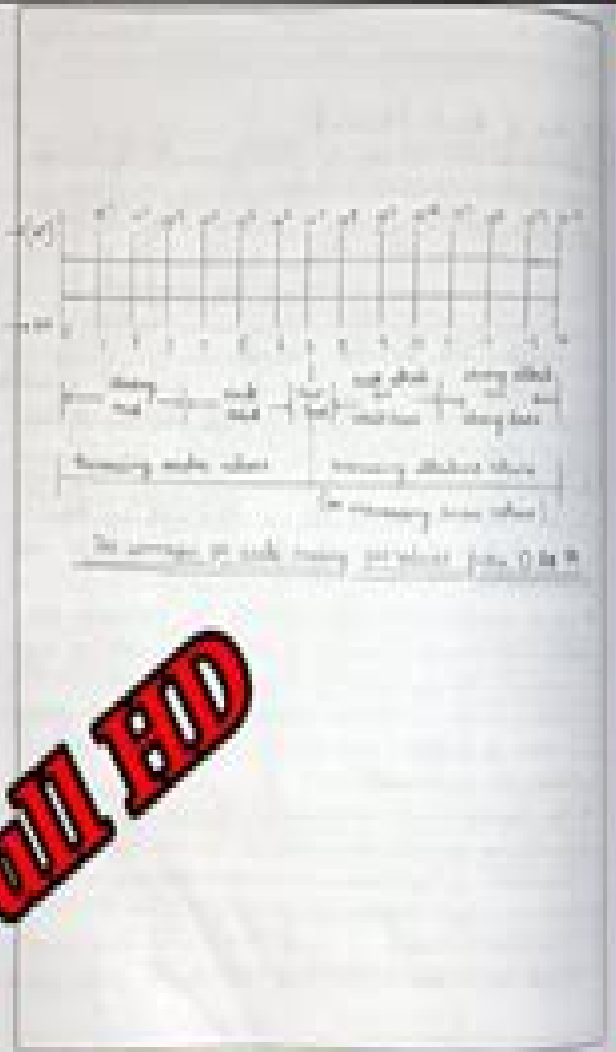
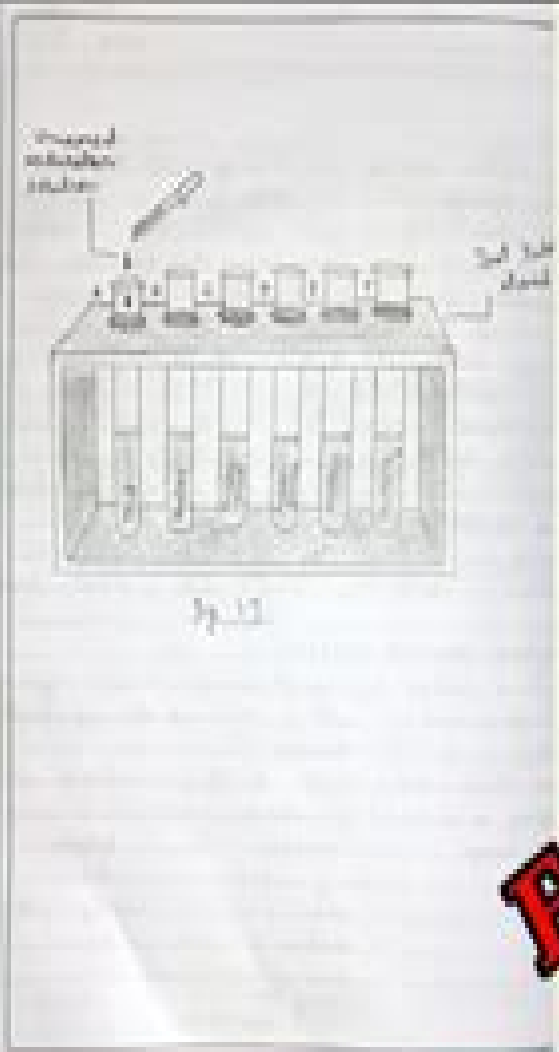
2. pH of a solution can be determined using a pH meter or using an indicator with universal indicator.

3. Universal Indicator Solution

It is a solution of several indicators having different pH ranges and it used to determine the approximate value of pH of a solution.

4. The colour that indicates the pH of a solution after adding a universal indicator is as follows:

pH range	Indicator	Colour
0-2	Strong acid	Red
3-4	Weak acid	Orange/Yellow
5-7	Neutral	Green
8-10	Weak base	Blue
11-14	Strong base	Dark Blue



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