

## CHAPTER - 2

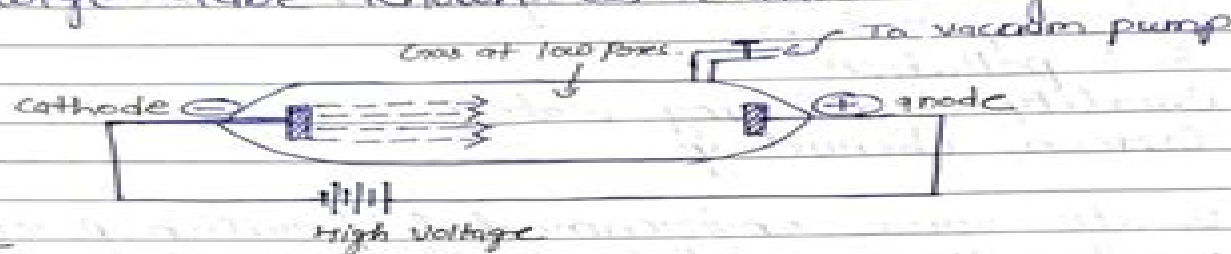
### STRUCTURE OF ATOM

#### Discovery of electron (Cathode Rays)

Cathode rays  $\rightarrow$  William Crookes (1879)

Electron ( $e^-$ )  $\rightarrow$  J. J. Thomson (-ve charge particle)

-ve charged rays which move from cathode to anode in discharge tube known as cathode rays.



#### properties

- 1 Cathode rays travel in a straight line (shadow)
- 2 Cathode rays carry -ve charge (electrostatic field deflection experiment)
- 3 Cathode rays are deflected by magnetic field
- 4 They are made up of material particles (rotary paddle wheel)
- 5 They produce heating effect & possess penetrating effect
- 6 They produce green fluorescence & X-rays (when they strike)
- 7 They also affect photographic plate.

#### Discovery of proton (Anode / Canal Rays)

Anode rays  $\rightarrow$  Goldstein (1886)

Proton ( $H^+$ )  $\rightarrow$  Goldstein & named by Rutherford (+ve charge particle)

+ve charged particle which move opposite the cathode in discharge tube known as anode rays.

# Class 11th Notes Chemistry Chapter Wise

**Patrick Vollmar**



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