

## Stoichiometric Calculations: Identify a compound using gravimetric analysis

Question 1: Atoms and molecules are way too small to be seen, even with a microscope, so how can we count how many there are?

- ✓ You count them by mass
- You count them using an electron microscope
- You can't count atoms and molecules
- You count them by volume

Question 2: In which of the pictures do you find the most particles? By particles I mean, sand grains, water molecules and iron atoms.

- ✓ There are roughly the same amount of particles on each picture
- In the iron bolt! It is quite heavy, and iron atoms are very small
- In the water glass! A shot of water isn't very large, but water molecules are super tiny
- In the Sahara desert! The desert is huge, and sand grains are small

Question 3: Why is it important to use demineralized water?

- ✓ Because demineralized water doesn't contain minerals
- Because demineralized water has a higher boiling point than regular tap-water
- Because demineralized water poses less of a health hazard in chemical laboratories
- Because demineralized water has a lower density than regular tap-water

Question 4: In order for the gravimetric analysis to be valid, we want all of the chloride ions to precipitate. How can we ensure this?

- ✓ By ensuring silver ions are in excess
- By adding the same volume of reactants
- By having roughly the same amount of each reactant
- By ensuring chloride ions are in excess

Question 5: Look at the reaction schemes in your journal or click the "View image" button below. How does the number of moles of  $\text{AgNO}_3$  silver nitrate correlate to the number of moles of the alkaline earth metal chloride?

- ✓ There should be twice as many moles of silver nitrate  $\text{AgNO}_3$  as unknown chloride  $\text{XCl}_2$
- We cannot tell before we have performed the reaction
- There should be half as many moles of silver nitrate  $\text{AgNO}_3$  as unknown chloride  $\text{XCl}_2$
- There should be the same number of moles of each

Question 6: How many moles of  $\text{AgNO}_3$  silver nitrate should we use to be sure that we have excess, no matter which of the three compounds it is?

- ✓ 0.100 mol
- 0.075 mol
- 0.050 mol
- 0.025 mol

Question 7: Look at your lab journal: Can you see from the reaction stoichiometry how the number of moles of  $\text{AgCl}$  silver chloride formed correlates to the number of moles of  $\text{AgNO}_3$  silver nitrate consumed?

- ✓ The number of moles are the same
- It differs in the three different reactions
- There are half as many moles of  $\text{AgCl}$  formed
- There are twice as many moles of  $\text{AgCl}$  formed

# Multiple Choice Questions On Gravimetric Analysis Calculations

**Mary Millhollon, Richard H. Langley**

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## **Table of Contents Multiple Choice Questions On Gravimetric Analysis Calculations**

1. Understanding the eBook Multiple Choice Questions On Gravimetric Analysis Calculations
  - The Rise of Digital Reading Multiple Choice Questions On Gravimetric Analysis Calculations
  - Advantages of eBooks Over Traditional Books
2. Identifying Multiple Choice Questions On Gravimetric Analysis Calculations
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Multiple Choice Questions On Gravimetric Analysis Calculations
  - User-Friendly Interface
4. Exploring eBook Recommendations from Multiple Choice Questions On Gravimetric Analysis Calculations
  - Personalized Recommendations
  - Multiple Choice Questions On Gravimetric Analysis Calculations User Reviews and Ratings
  - Multiple Choice Questions On Gravimetric Analysis Calculations and Bestseller Lists

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  - Multiple Choice Questions On Gravimetric Analysis Calculations eBook Subscription Services
  - Multiple Choice Questions On Gravimetric Analysis Calculations Budget-Friendly Options
6. Navigating Multiple Choice Questions On Gravimetric Analysis Calculations eBook Formats
  - ePub, PDF, MOBI, and More
  - Multiple Choice Questions On Gravimetric Analysis Calculations Compatibility with Devices
  - Multiple Choice Questions On Gravimetric Analysis Calculations Enhanced eBook Features
7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Multiple Choice Questions On Gravimetric Analysis Calculations
  - Highlighting and Note-Taking Multiple Choice Questions On Gravimetric Analysis Calculations
  - Interactive Elements Multiple Choice Questions On Gravimetric Analysis Calculations
8. Staying Engaged with Multiple Choice Questions On Gravimetric Analysis Calculations
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Multiple Choice Questions On Gravimetric Analysis Calculations
9. Balancing eBooks and Physical Books Multiple Choice Questions On Gravimetric Analysis Calculations
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Multiple Choice Questions On Gravimetric Analysis Calculations
10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
11. Cultivating a Reading Routine Multiple Choice Questions On Gravimetric Analysis Calculations
  - Setting Reading Goals Multiple Choice Questions On Gravimetric Analysis Calculations
  - Carving Out Dedicated Reading Time
12. Sourcing Reliable Information of Multiple Choice Questions On Gravimetric Analysis Calculations
  - Fact-Checking eBook Content of Multiple Choice Questions On Gravimetric Analysis Calculations
  - Distinguishing Credible Sources
13. Promoting Lifelong Learning

- Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
14. Embracing eBook Trends
- Integration of Multimedia Elements
  - Interactive and Gamified eBooks

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