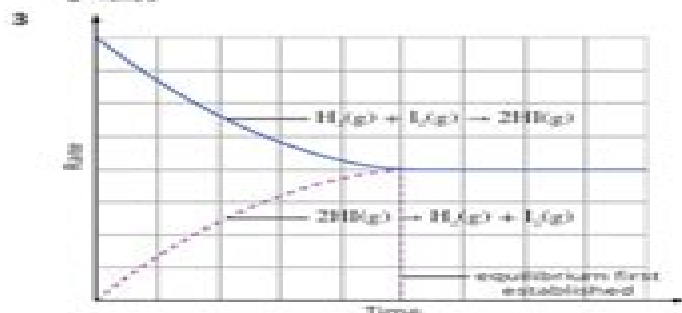


## Chapter 2 Chemical equilibrium

### 2.1 Dynamic equilibrium

#### 2.1 KEY QUESTIONS

- An **open** system occurs when both matter and energy are exchanged between the system and the surroundings.
  - Chemical equilibria can occur in **closed** systems.
  - In some circumstances, a reversible reaction can form a **dynamic** equilibrium.
  - The extent of reaction for chemical equilibria is **different** for different equilibria.
- false
  - false
  - true
  - false



- closed, decreases, less, reverse, increases, equal
- open; both energy and matter (carbon dioxide and water vapour) are exchanged with the surroundings
  - closed; only energy can be exchanged with the surroundings. While the can is unopened, no matter is exchanged.
  - open; both energy and matter are exchanged with the surroundings
- B and F. Equilibrium can only occur in a closed system. At the molecular level, some of the sugar in the saturated sugar solution will continue to dissolve at the same rate as some is precipitated. In the unopened bottle of sparkling mineral water, the dissolved carbon dioxide will be in equilibrium with the gaseous carbon dioxide. In all of the other situations, the reverse reaction is not occurring at the same rate as the forward reaction, because the systems are open.
- $130\text{ kJ mol}^{-1}$ . The activation energy is the difference between  $390\text{ kJ mol}^{-1}$  and  $260\text{ kJ mol}^{-1}$ , which is  $130\text{ kJ mol}^{-1}$ .
- Chemical equilibrium is 'dynamic' because both forward and reverse reactions occur at the same rate. An equilibrium develops between water vapour and water when wet clothes are in a sealed bag, with water evaporating as rapidly as water vapour condenses, so the clothes remain wet.
  - When the bag is opened, water vapour escapes and the rate of evaporation exceeds the rate of condensation. The system is not in equilibrium and the clothes dry.
- $0.07\text{ mol L}^{-1}$
  - $0\text{ mol L}^{-1}$
  - $0.03\text{ mol L}^{-1}$
  - $0.08\text{ mol L}^{-1}$
  - $0.04\text{ moles}$
  - The horizontal regions of the graph indicate that there is no change in concentration of  $\text{NO}_2$  or  $\text{N}_2\text{O}_4$ . This is when the system is at equilibrium and both the forward and reverse reactions are occurring at the same rate.
  - 6 seconds
  - As the reaction proceeds, the intensity of the brown colour will increase as the concentration of  $\text{NO}_2$  increases until equilibrium is reached. Once equilibrium is reached, the colour of the reaction mixture will remain constant.

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**Y Pai**



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## **Table of Contents Pearson Reaction Rates And Equilibrium Packet Answers**

1. Understanding the eBook Pearson Reaction Rates And Equilibrium Packet Answers
  - The Rise of Digital Reading Pearson Reaction Rates And Equilibrium Packet Answers
  - Advantages of eBooks Over Traditional Books
2. Identifying Pearson Reaction Rates And Equilibrium Packet Answers
  - Exploring Different Genres
  - Considering Fiction vs. Non-Fiction
  - Determining Your Reading Goals
3. Choosing the Right eBook Platform
  - Popular eBook Platforms
  - Features to Look for in an Pearson Reaction Rates And Equilibrium Packet Answers
  - User-Friendly Interface
4. Exploring eBook Recommendations from Pearson Reaction Rates And Equilibrium Packet Answers
  - Personalized Recommendations
  - Pearson Reaction Rates And Equilibrium Packet Answers User Reviews and Ratings
  - Pearson Reaction Rates And Equilibrium Packet Answers and Bestseller Lists
5. Accessing Pearson Reaction Rates And Equilibrium Packet Answers Free and Paid eBooks
  - Pearson Reaction Rates And Equilibrium Packet Answers Public Domain eBooks
  - Pearson Reaction Rates And Equilibrium Packet Answers eBook Subscription Services

- Pearson Reaction Rates And Equilibrium Packet Answers Budget-Friendly Options
- 6. Navigating Pearson Reaction Rates And Equilibrium Packet Answers eBook Formats
  - ePub, PDF, MOBI, and More
  - Pearson Reaction Rates And Equilibrium Packet Answers Compatibility with Devices
  - Pearson Reaction Rates And Equilibrium Packet Answers Enhanced eBook Features
- 7. Enhancing Your Reading Experience
  - Adjustable Fonts and Text Sizes of Pearson Reaction Rates And Equilibrium Packet Answers
  - Highlighting and Note-Taking Pearson Reaction Rates And Equilibrium Packet Answers
  - Interactive Elements Pearson Reaction Rates And Equilibrium Packet Answers
- 8. Staying Engaged with Pearson Reaction Rates And Equilibrium Packet Answers
  - Joining Online Reading Communities
  - Participating in Virtual Book Clubs
  - Following Authors and Publishers Pearson Reaction Rates And Equilibrium Packet Answers
- 9. Balancing eBooks and Physical Books Pearson Reaction Rates And Equilibrium Packet Answers
  - Benefits of a Digital Library
  - Creating a Diverse Reading Collection Pearson Reaction Rates And Equilibrium Packet Answers
- 10. Overcoming Reading Challenges
  - Dealing with Digital Eye Strain
  - Minimizing Distractions
  - Managing Screen Time
- 11. Cultivating a Reading Routine Pearson Reaction Rates And Equilibrium Packet Answers
  - Setting Reading Goals Pearson Reaction Rates And Equilibrium Packet Answers
  - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Pearson Reaction Rates And Equilibrium Packet Answers
  - Fact-Checking eBook Content of Pearson Reaction Rates And Equilibrium Packet Answers
  - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
  - Utilizing eBooks for Skill Development
  - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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