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REPORT SHEET | LAB
**Compounds and
 Their Bonds** | **9**

A. Ions: Transfer of Electrons

Element	Atomic Number	1. Electron Configuration of Atom	2. Electron-Dot Symbol	3. Loss or Gain of Electrons	4. Electron Configuration of Ion	5. Ionic Charge	6. Symbol of Ion	7. Name of Ion
Sodium	11	$1s^2 2s^2 2p^6 3s^1$	Na [•]	lose 1 e ⁻	$1s^2 2s^2 2p^6$	1+	Na ⁺	Sodium
Nitrogen	7	$1s^2 2s^2 2p^3$	$:\ddot{N}:$	gains 3e ⁻	$1s^2 2s^2 2p^6$	3-	N ³⁻	nitride
Aluminum	13	$1s^2 2s^2 2p^6 3s^2 3p^1$	$\cdot\ddot{Al}\cdot$	lose 3e ⁻	$1s^2 2s^2 2p^6$	3+	Al ³⁺	aluminum
Chlorine	17	$1s^2 2s^2 2p^6 3s^2 3p^5$	$:\ddot{Cl}:$	gains 1e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	1-	Cl ⁻	chloride
Calcium	20	$1s^2 2s^2 2p^6 3s^2 3p^6 4s^2$	Ca ^{••}	lose 2e ⁻	$1s^2 2s^2 2p^6 3s^2 3p^6$	2+	Ca ²⁺	calcium
Oxygen	8	$1s^2 2s^2 2p^4$	$:\ddot{O}:$	gains 2e ⁻	$1s^2 2s^2 2p^6$	2-	O ²⁻	oxide

* $1s^2 2s^2 2p^6 3s^2 3p^6$

B. Ionic Compounds and Formulas

1. Physical properties

Compound	Appearance	Density	Melting Point
Sodium chloride, NaCl	white crystalline solid	2.17 g/cm ³	1,474 °F

2. Formulas of ionic compounds

Name	Positive Ion	Negative Ion	Formula
Lithium iodide	Li ⁺	I ⁻	LiI
Aluminum oxide	Al ³⁺	O ²⁻	Al ₂ O ₃
Calcium sulfide	Ca ²⁺	S ²⁻	CaS
Magnesium bromide	Mg ²⁺	Br ⁻	MgBr ₂
Potassium nitride	K ⁺	N ³⁻	K ₃ N
Sodium fluoride	Na ⁺	F ⁻	NaF

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