



3: Chemical Kinetics

Standard: 12th Science

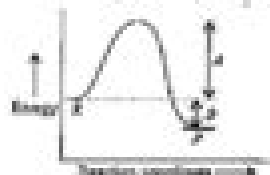
Subject: Chemistry

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Q1. Consider the graph given in Fig. Which of the following options does not show instantaneous rate of reaction at 40th second?

1. $\frac{V_2 - V_1}{50 - 30}$
2. $\frac{V_1 - V_2}{50 - 30}$
3. $\frac{V_2 - V_1}{30 - 20}$
4. $\frac{V_1 - V_2}{30 - 20}$

Q2. The energy diagram of a reaction $P + Q \rightarrow R + S$ is given. What are A and B in the graph?



1. A → activation energy, B → heat of reaction
 2. A → threshold energy, B → heat of reaction
 3. A → heat of energy, B → activation reaction
 4. A → potential energy, B → energy of reaction
- Q3.** Decomposition of nitrogen pentoxide is known to be a first order reaction 75 percent of the oxide had decomposed in the first 24 minutes. At the end of an hour, after the start of the reaction, the amount of oxide left will be
1. Nil
 2. About 1%
 3. About 2%
 4. About 3%
- Q4.** Higher order (>3) reactions are rare due to:
1. Low probability of simultaneous collision of all the reacting species.
 2. Increase in entropy and activation energy as more molecules are involved.



Practice Chemical Kinetics Questions Answer

Rachel S Tattersall



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Table of Contents Practice Chemical Kinetics Questions Answer

1. Understanding the eBook Practice Chemical Kinetics Questions Answer
 - The Rise of Digital Reading Practice Chemical Kinetics Questions Answer
 - Advantages of eBooks Over Traditional Books
2. Identifying Practice Chemical Kinetics Questions Answer
 - Exploring Different Genres
 - Considering Fiction vs. Non-Fiction
 - Determining Your Reading Goals
3. Choosing the Right eBook Platform
 - Popular eBook Platforms
 - Features to Look for in an Practice Chemical Kinetics Questions Answer
 - User-Friendly Interface
4. Exploring eBook Recommendations from Practice Chemical Kinetics Questions Answer
 - Personalized Recommendations
 - Practice Chemical Kinetics Questions Answer User Reviews and Ratings
 - Practice Chemical Kinetics Questions Answer and Bestseller Lists
5. Accessing Practice Chemical Kinetics Questions Answer Free and Paid eBooks
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 - Practice Chemical Kinetics Questions Answer eBook Subscription Services

- Practice Chemical Kinetics Questions Answer Budget-Friendly Options
- 6. Navigating Practice Chemical Kinetics Questions Answer eBook Formats
 - ePub, PDF, MOBI, and More
 - Practice Chemical Kinetics Questions Answer Compatibility with Devices
 - Practice Chemical Kinetics Questions Answer Enhanced eBook Features
- 7. Enhancing Your Reading Experience
 - Adjustable Fonts and Text Sizes of Practice Chemical Kinetics Questions Answer
 - Highlighting and Note-Taking Practice Chemical Kinetics Questions Answer
 - Interactive Elements Practice Chemical Kinetics Questions Answer
- 8. Staying Engaged with Practice Chemical Kinetics Questions Answer
 - Joining Online Reading Communities
 - Participating in Virtual Book Clubs
 - Following Authors and Publishers Practice Chemical Kinetics Questions Answer
- 9. Balancing eBooks and Physical Books Practice Chemical Kinetics Questions Answer
 - Benefits of a Digital Library
 - Creating a Diverse Reading Collection Practice Chemical Kinetics Questions Answer
- 10. Overcoming Reading Challenges
 - Dealing with Digital Eye Strain
 - Minimizing Distractions
 - Managing Screen Time
- 11. Cultivating a Reading Routine Practice Chemical Kinetics Questions Answer
 - Setting Reading Goals Practice Chemical Kinetics Questions Answer
 - Carving Out Dedicated Reading Time
- 12. Sourcing Reliable Information of Practice Chemical Kinetics Questions Answer
 - Fact-Checking eBook Content of Practice Chemical Kinetics Questions Answer
 - Distinguishing Credible Sources
- 13. Promoting Lifelong Learning
 - Utilizing eBooks for Skill Development
 - Exploring Educational eBooks
- 14. Embracing eBook Trends

- Integration of Multimedia Elements
- Interactive and Gamified eBooks

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